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In a full glass of water, the water surface seems to bulge over the rim of the glass. Water forms nearly spherical drops at the end of an eyedropper.

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The hydrogen bonding in water cause the high surface tension and low vapor pressure of water. After you claim an answer you'll have 24 hours to send in a draft. An editor will review the submission and either publish your submission or provide feedback. Next Answer Chapter 15 - Water and Aqueous

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A water molecule, because of its bent shape, is a polar molecule. After you claim an answer you'll have 24 hours to send in a draft. An editor will review the submission and either publish your submission or provide feedback. Next Answer Chapter 15 - Water and Aqueous Systems - 15 Assessment

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### **SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)**

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CH<sub>4</sub> doesn't dissolve in water because it is a molecular compound with no net dipole KCl is soluble because of its ionic bonds. The hydrogen and oxygen break the molecules apart because of the different charges of both KCl and H<sub>2</sub>O.

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Hydrogen ion ( $\text{H}_3\text{O}^+$ ) as a partially flattened pyramid . Hydrogen ions .

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Water is a simple tri-atomic molecule,  $\text{H}_2\text{O}$ ; Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values)

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